Additions and Corrections

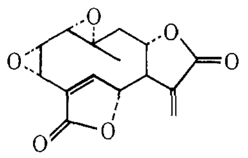
Vol. 35, 1970

F. R. Stermitz and F. A. Norris: Oxidative Acylation. A New Reaction of Primary Nitro Compounds.

Page 527. It has come to our attention that T. Urbanski and W. Guzynaska, *Roczniki Chem.*, **25**, 213 (1951), reported the isolation in 12% yield of an unknown compound by treating nitropropane with ketene and NaOAc at 100°. The unknown was thought to be either $CH_3CH(N{=}O)OAc$ or *N,O*-diacetyl-*N*-propionylhydroxylamine. We reported the latter as a new compound (structure IIa) in the above paper. It is apparent from the method of preparation that Urbanski's unknown compound was indeed *N,O*-diactyl-*N*-propionylhydroxylamine and that he had achieved the first oxidative acylation, albeit in a less practical sense (compare our 70% yield of IIa from nitropropane by warming with acetic anhydride).

W. Herz, P. S. Subramaniam, P. S. Santhanam, K. Aota, and A. L. Hall: Structure Elucidation of Sesquiterpene Lactones from *Mikania scandens* (L.) Willd.

Page 1454. Column 1, first structure. Formula 1 should be



Samuel P. McManus, John T. Carroll, and Charles U. Pittman, Jr.: Acid-Catalyzed Cyclization Reactions. IX. The Formation of Oxazolinium and Thiazolinium Cations from *N*-Allyl- and Substituted *N*-Allylamides, -urethanes, -ureas, and -thioureas.

Page 3770. The chemical shifts of the ring protons in thiazolinium ions **2p** and **2q** were incorrectly assigned. In Table II, the correct values for cation **2p** are δ 4.30 (m) and 4.70 (m, C-4 protons) and 4.70 (m, C-5 protons). For cation **2q**, the correct values are δ 4.19 (m) and 4.60 (m, C-4 protons) and 4.60 (m, C-5 protons). All peaks are broad and, in each case, one C-4 and one C-5 proton are not resolved.

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D. E. Dorman, M. Jautelat and J. D. Roberts*: Carbon-13 Nuclear Magnetic Resonance Spectroscopy. Quantitative Correlations of the Carbon Chemical Shifts of Acyclic Alkenes.

We are indebted to Dr. V. D. Mochel for pointing out inconsistencies in the parameters reported for the calculation of ^{13}C chemical shifts of alkenes which have turned out to be proofreading errors.

Page 2761. In Table III, footnote *b*, the value of $(\alpha + \alpha')$ should be -2.61 ± 0.12 ppm.

Page 2761. In Table IV, the correction parameter for cis in the first column should be 0.55 ± 0.02 ppm. A line was omitted from this table which follows directly after the cis parameter as follows.

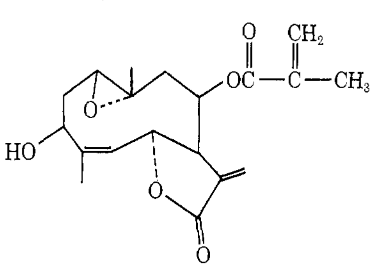
Corr β 2.82 ± 0.08 1.832 ± 0.019 2.686 ± 0.012

Page 2763. The last two sentences of the first paragraph of section B were actually inoperative for the resonances considered in Table VII.

Vol. 37, 1972

W. Herz* and S. V. Bhat: Woodhousin, a New Germacranolide from *Bahia woodhousei* (Gray) Gray.

Page 910. Column 1, formula **15**. Formula **15** should be



M. J. Tremelling and J. M. McBride*: Solvent Steric Effects. V. Azobis-2-methyl-3-phenyl-2-butane. The Absolute Configuration of Some Derivatives of 2-Methyl-3-phenylbutane.

Page 1073. Professor Luciano Lardicci (University of Pisa) has called to our attention that the original assignment of the *S* absolute configuration to (–)-2-methyl-3-phenylbutane (II) by Červinka and Hub is incorrect [*cf.* O. Červinka, V. Dudek, and L. Hub, *Z. Chem.*, 9, 267 (1969); L. Lardicci and R. Menicagli, *Chim. Ind.* (*Milan*), **51**, 1387 (1969); and D. R. Clark and H. S. Mosher, *J. Org. Chem.*, **35**, 1114 (1970)]. Since all our configurational assignments are based on correlation with II, each one should be reversed.

L. A. Hulshof, Aafje Vos,* and Hans Wynberg: The Crystal and Molecular Structure and Absolute Configuration of *d*-Spiro-[3.3]heptane-2,6-dicarboxylic Acid at –160°.

Page 1767. We reported that the dextrorotatory Fecht acid or *d*-spiro[3.3]heptane-2,6-dicarboxylic acid (see Figure 1, page 1768) has the *R* configuration. The assignment was based on (a) application of Lowe's rule to the spiro[3.3]heptane system, (b) special solvent effects of *d*-Fecht acid, (c) use of Klyne's sector rule for carboxylic acids and esters, and (d) X-ray diffraction with anomalous scattering. Although the last-named method, *viz.*, the anomalous X-ray scattering studies, could in principle give definitive information, this was not the case with Fecht acid because of the very small differences in anomalous scattering between the carbon and oxygen atoms. Thus, as stated in our previous paper, only eight Bijvoet pairs (five with Cr Kα radiation and three with Cu Kα radiation) had intensity differences of sufficient magnitude to be considered and only a 75% correspondence between observed and calculated intensity order for hkl and $\bar{h}\bar{k}\bar{l}$ reflections was obtained. The X-ray work was therefore repeated using the dextrorotatory barium salt of *d*-Fecht acid with the composition $C_9H_{12}O_4 \cdot C_{18}H_{22}O_8Ba \cdot H_2O$. The anomalous scattering effect of the barium atom was, as expected, sufficiently large to allow an unambiguous assignment of the absolute configuration of Fecht acid. So, in contrast to the assignment suggested previously we now find that *d*-Fecht acid has the *S* instead of the *R* configuration (see the correct absolute configuration of *d*-Fecht acid in the Figure 1′). A full account of the work will be given (B. van Dijk, L. A. Hulshof, H. Wynberg, and J. L. de Boer, manuscript to be submitted to *J. Amer. Chem. Soc.*) and an attempt will be made to explain the failure of the methods previously used for the determination of the absolute configuration of *d*-Fecht acid. Preliminary studies suggest that the nonplanarity of the four-membered ring is a major factor causing the discrepancies between theory and experiment.

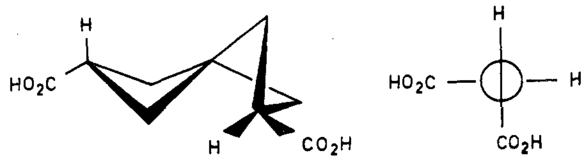


Figure 1′.—*S*-(+)-Spiro[3.3]heptane-2,6-dicarboxylic acid with the notation of R. S. Cahn, C. K. Ingold, and V. Prelog, *Experientia*, **12**, 81 (1956); *Angew. Chem.*, **78**, 417 (1966).